

188660 WILLIAM SANATA SANATA PERSEASA SANATA

OFFICE OF NAVAL RESEARCH

Contract N00014-K-0043

R & T Code 413f0001---01

TECHNICAL REPORT No. 34

Vibrational Motions of Buckminsterfullerene

by

Z. C. Wu, Daniel Jelski and Thomas F. George

Prepared for Publication

in

Chemical Physics Letters

Departments of Chemistry and Physics State University of New York at Buffalo Buffalo, New York 14260

March 1987

Reproduction in whole or in part is permitted for any purpose of the United States Government.

This document has been approved for public release and sale; its distribution is unlimited.



•	•	^			•	٧	-		•	o	٠.	•	^			٠.		-	 •	•		44		•	•		2
-		•	u		1	▼ .	٠.	_	•	3	ш	•	L	-		,,,,	•		•		м		•	•		9 8	z

REPORT DOCUMENTATION PAGE										
1a REPORT SECURITY CLASSIFICATION Unclassified		1b. RESTRICTIVE MARKINGS								
24. SECURITY CLASSIFICATION AUTHORITY		3. DISTRIBUTION/AVAILABILITY OF REPORT Approved for public release; distribution								
26. DECLASSIFICATION/DOWNGRADING SCHED	PULE	unlimited								
4. PERFORMING ORGANIZATION REPORT NUM	BER(S)	5. MONITORING ORGANIZATION REPORT NUMBER(S)								
UBUFFALO/DC/87/TR-34	E									
Depts. Chemistry & Physics State University of New York	6b. OFFICE SYMBOL (If applicable)	76. NAME OF MONITORING ORGANIZATION								
6c. ADDRESS (City, State and ZIP Code)		76. ADDRESS (City.		le)						
Fronczak Hall, Amherst Campus		Chemistry P			j					
Buffalo, New York 14260		800 N. Quin	cy Street							
		Arlington,	Virginia 22	2217						
Be. NAME OF FUNDING/SPONSORING ORGANIZATION	8b. OFFICE SYMBOL (If applicable)	9. PROCUREMENT I	NSTRUMENT ID	ENTIFICATION N	JMBER					
Office of Naval Research	<u> </u>									
8c. ADDRESS (City, State and ZIF Code)	!	10. SOURCE OF FUN	1		·					
Chemistry Program		PROGRAM ELEMENT NO.	PROJECT NO.	TASK NO.	WORK UNIT					
800 N. Quincy Street			17.	,,,,,	"					
Arlington, Virginia 22217		l	<u> </u>	<u>L</u>	J					
Vibrational Motic	ons of Buckminst	erfullerene —————			_					
	aniel Jelski an			100 000 0						
13a TYPE OF REPORT 13b. TIME C	OVERED TO	14. DATE OF REPORT (Yr., Mo., Day) 15. PAGE COUNT March 1987 14								
16. SUPPLEMENTARY NOTATION		1001011	150,							
Prepared for Publication	n in Chemical Ph	ysics Letters								
17. COSATI CODES	18. SUBJECT TERMS (C	ontinue on reverse if ne	cessary and identi	ify by block number	1					
FIELD GROUP SUB. GR.	BUCKMINSTERFULL	LERENE NON-CARTESIAN COORDINATES								
	CARBON (60-ATOM			UR FORCE CON						
	COMPLETE VIBRAT		M < 180	D BY 180 MAT	RIX					
19. ASSTRACT (Continue on reverse if necessary and										
A non-Cartesian co	pordinate system	is developed	which perm	nits the						
vibrational motions of										
in terms of four force constants. A 180 x 180 matrix is then derived which,										
when diagonalized, yields the complete vibrational spectrum. These results										
are compared with those obtained previously via a MNDO calculation.										
20. DISTRIBUTION/AVAILABILITY OF ABSTRAC	СТ	21. ABSTRACT SECU	JRITY CLASSIFI	CATION						
UNCLASSIFIED/UNLIMITED 🖾 SAME AS RPT.	TIC USERS	Unclassifi	ed							
22L NAME OF RESPONSIBLE INDIVIDUAL		226. TELEPHONE N		22c. OFFICE SYM	BOL					
Dr. David L. Nelso	n	(Include Area Co (202) 696-								

VIBRATIONAL MOTIONS OF BUCKMINSTERFULLERENE

Z. C. Wu, Daniel A. Jelski and Thomas F. George

Departments of Physics & Astronomy and Chemistry 239 Fronczak Hall State University of New York at Buffalo Buffalo, New York 14260

Abstract

A non-Cartesian coordinate system is developed which permits the vibrational motions of Buckminsterfullerene (Bucky ball) to be expressed in terms of four force constants. A 180×180 matrix is then derived which, when diagonalized, yields the complete vibrational spectrum. These results are compared with those obtained previously via a MNDO calculation.



Accesi	on For	7						
DTIC	ounced							
By								
Availability Codes								
Dist	Avail and Specia							
A-1								

I. <u>Introduction</u>

The purpose of this paper is to investigate the vibrational motions of Buckminsterfullerene (Bucky ball) in a way that lends itself toward future application. Hence simplicity, both theoretical and practical, is an important consideration. To this end we can take advantage of the symmetry of the species, and we shall derive a relatively simple procedure for calculating the vibrational normal frequencies. Furthermore, our method lends itself to physical interpretation, and we shall be able to make some statements about the nature of Bucky ball. The experimental literature on Bucky Ball is well known. Theoretically much work has been dedicated to determining the most stable structure of C_{60} clusters, and to determining some of their optical properties. We continue this work here by assuming that Bucky ball exists and is relatively stable.

In essence our calculation can be described as follows. We express the position of each atom in local coordinates, choosing our coordinate system (non-Cartesian) in a way that simplifies the calculation. This permits us to express the equation of motion of each atom in terms of four force constants, and since all atoms are identical we immediately derive an equation of motion for the entire system. The fundamental principles behind this calculation are found in Ref. 12. While we must input the force constants into our calculation (this is a disadvantage), we generate the correct number of modes, the correct point group properties, and, at least qualitatively, the correct frequencies. This requires only the diagonalization of one 180 × 180 matrix, a procedure which requires only about three minutes CPU on a VAX. Furthermore, since we can vary the force

constants and since we can compare our results with those of Stanton, ¹³ we are able to make some predictions about the structure of Bucky ball as compared to benzene. In the next section we derive our method for calculating the normal mode frequencies, and in Section III we compare our results with those of Stanton ¹³ and discuss the implications.

II. Theory

The structure of Buckminsterfullerene, illustrated in Fig. 1, is that of a truncated icosahedron with sixty vertices, twenty hexagonal faces and twelve pentagonal faces. A carbon atom occupies each vertex. The bonds separating a hexagon from a pentagon are found to be more "single" than bonds separating two hexagons. 6,8 The bond lengths are given as 1.41 Å and 1.54 Å, respectively. 8 Since each vertex is at the intersection of two hexagonal and one pentagonal face, it follows that it is joined by one "double" bond and two "single" bonds. This implies that the motion of an atom around it equilibrium position can be decomposed along the three noncoplanar directions of the adjacent bonds (Fig. 2). We can use x_i , y_i and \boldsymbol{z}_{i} to denote the deviation of atom i from its equilibrium position in these directions (we caution the reader not to confuse our notation with Cartesian coordinates, since ours are not Cartesian). The angle between two adjacent single bonds is $\frac{3\pi}{5}$ and that between a double bond and an adjacent single bond is $\frac{2\pi}{3}$. One can work out the angle between a double bond and its adjacent pentagonal face as $\Psi = \cos^{-1}(\frac{1}{2\cos\frac{3\pi}{\epsilon}})$. The angle between two adjacent hexagonal faces is $\phi = \cos^{-1} \left(\frac{8}{3} (\sin \frac{3\pi}{10})^2 - 1 \right)$, and the angle between adjacent hexagonal and pentagonal faces is

$$\theta = \frac{\pi - \phi}{2} - \Psi.$$

The variations of the bond length a_{ij} , b_{ik} and c_{il} (see Fig. 2) due to the motion of the atoms are, to first order,

$$\delta a_{ij} = (x_i - \frac{1}{2}y_i - \frac{1}{2}z_i) + (x_j - \frac{1}{2}y_j - \frac{1}{2}z_j),$$

$$\delta b_{ik} = (y_i - \frac{1}{2}x_i - \cos\frac{2\pi}{5}z_i) + (z_k - \frac{1}{2}x_k - \cos\frac{2\pi}{5}y_k),$$

$$\delta c_{ik} = (z_i - \frac{1}{2}x_i - \cos\frac{2\pi}{5}y_i) + (y_k - \frac{1}{2}x_k - \cos\frac{2\pi}{5}z_k).$$
(1)

Also, the variations of the angles between bonds α , β and γ are

$$\delta\alpha_{i} = \frac{1}{L}(-\sin\frac{\pi}{3}\cos\theta \ y_{j} + \sin\frac{\pi}{3} \ z_{j} + \sin\frac{\pi}{3} \ x_{k} - \sin\frac{2\pi}{5}\cos\theta \ y_{k} \\ - \sin\frac{\pi}{3} \ y_{i} + \sin\frac{\pi}{3}\cos\theta \ z_{i} - \sin\frac{\pi}{3} \ x_{i} + \sin\frac{2\pi}{5}\cos\theta \ z_{i}),$$

$$\delta\beta_{i} = \frac{1}{L}(-\sin\frac{\pi}{3}\cos\theta \ z_{j} + \sin\frac{\pi}{3} \ y_{j} + \sin\frac{\pi}{3} \ x_{\ell} - \sin\frac{2\pi}{5}\cos\theta \ z_{\ell} \\ - \sin\frac{\pi}{3} \ z_{i} + \sin\frac{\pi}{3}\cos\theta \ y_{i} - \sin\frac{\pi}{3} \ x_{i} + \sin\frac{2\pi}{5}\cos\theta \ y_{i}), \quad (2)$$

$$\delta \gamma_{i} = \frac{1}{L} (\sin \frac{2\pi}{5} y_{k} - \sin \frac{\pi}{3} \cos \theta x_{k} + \sin \frac{2\pi}{5} z_{\ell} - \sin \frac{\pi}{3} \cos \theta x_{\ell} + \sin \frac{\pi}{3} \cos \theta x_{i} - \sin \frac{2\pi}{5} y_{i} + \sin \frac{\pi}{3} \cos \theta x_{i} - \sin \frac{2\pi}{5} z_{i}),$$

where L refers to the bond lengths which are assumed equal.

The kinetic energy of each atom is

$$T_{i} = \frac{m}{2} \left[\left(\dot{x}_{i} - \frac{1}{2} \dot{y}_{i} - \frac{1}{2} \dot{z}_{i} \right)^{2} + \left(\dot{y}_{i} \cos \frac{\pi}{5} - \dot{z}_{i} \cos \frac{\pi}{5} \right)^{2} + \left(\dot{y}_{i} \cos \frac{3\pi}{10} \sin \Psi + \dot{z}_{i} \cos \frac{3\pi}{10} \sin \Psi \right)^{2} \right], \tag{3}$$

where m is the mass of each atom. It is assumed that the atoms undergo harmonic oscillations around their equilibrium positions, where the potential is due to the variations of the bond lengths and the angles between them. If atoms j, k and 1 represent nearest neighbors of atom i, as shown in Fig. 2, then the Lagrangian for this system takes the form

$$L = T - V = \sum_{i} T_{i} - \sum_{i>j} \frac{1}{2} k_{1} \delta a_{ij}^{2} - \sum_{i>k} \frac{1}{2} k_{2} \delta b_{ik}^{2} - \sum_{i>k} \frac{1}{2} k_{2} \delta c_{ik}^{2}$$
$$- \sum_{i} \frac{1}{2} [k_{3} (\delta \alpha_{i}^{2} + \delta \beta_{i}^{2}) + k_{4} \delta \gamma_{i}^{2}], \qquad (4)$$

where k_1 , k_2 , k_3 and k_4 are all force constants. Equation (4) can be rewritten in the compact form

$$L = \frac{1}{2} \sum_{m_1, n=1}^{180} (T_{mn} X^m X^n - V_{mn} X^m X^n), \qquad (5)$$

and the appropriate Euler-Lagrange equation is

$$\ddot{X}^{n} = -\sum_{m, \, k=1}^{180} (T^{-1})^{mn} V_{mk} X^{k}. \tag{6}$$

By diagonalizing the 180×180 matrix $T^{-1}V$, we get the squares of the eigenfrequencies along with their associated eigenvectors.

Several cases are worth discussing:

- (a) $k_2 = k_3 = k_4 = 0$. Each double bond with force constant k_1 becomes independent. Then one obtains only one non-zero eigenfrequency, $(k_1/m)^{1/2}$, with a 30-fold degeneracy.
- (b) $k_1 = k_3 = k_4 = 0$. The system reduces to 12 independent pentagons. For each pentagon the eigenfrequencies are (with degeneracies) 0.89401 (1), 1.00029 (2) and 1.15166 (2), with $k_2/m = 1$. All other eigenfrequencies are
- (c) $k_3 = k_4 = 0$. One obtains 90 non-zero eigenfrequencies corresponding to the 90 bonds. Since there is no force on the angles, Bucky ball can be squished in 90 6 = 84 different ways. The remaining 6 degrees of freedom are due to translation and rotation.
- (d) In the general case, there are always 6 zero eigenfrequencies corresponding to global translational and rotational motions. There are also three non-degenerate modes.

III. Results and Discussion

Table I contains 180 eigenvalues representing the vibrational frequencies of Buckminsterfullerene. The values of \mathbf{k}_1 - \mathbf{k}_4 are those that best seem to fit the data of Stanton, 13 which was calculated by the MNDO method with the gradients evaluated at six points around an atom. The best fit was determined by matching the highest and lowest frequencies and also by matching the three non-degenerate frequencies. In addition, we have insisted that the ratio between \mathbf{k}_1 and \mathbf{k}_2 be approximately equal to the

inverse ratio of the bond lengths. Since at this point we have no independent way of determining k_4 , we have insisted that it equal k_3 .

For benzene the appropriate values would be $k_1 = k_2 = 7.62 \times 10^5$ dynes/cm and $k_3 = .667 \times 10^5$ dynes/cm. ¹⁴ Comparing these values with those given in Table I, we see that the force constants for Bucky ball are considerably higher than for benzene. This is physically reasonable since the stretching of a bond in Bucky ball involves a dislocation of the entire molecule, and hence greater resistance. Similarly we expect the stretching constant to be larger.

It is also instructive to look at the three non-degenerate modes. Stanton 13 has calculated their frequencies (in cm $^{-1}$) as 611 (for the Ag mode), 973 (Au) and 1667 (Ag). Our calculation reveals frequencies at 548, 970, and 1627 cm $^{-1}$, respectively. We have found that the two Ag modes depend only on constants k_1 and k_2 . The first is a "breathing" motion whereas the second involves oscillation between the "single" and "double" bonds. In both cases angles can be expected to play no role. The Au motion, on the other hand, involves a twisting (angle-dependent) motion of the pentagons. When k_3 and k_4 are zero, the eigenfrequency of this motion becomes zero. All other modes must be degenerate according to the point group of Buckminsterfullerene. 13

Our method is well suited to elucidating this sort of physical description of the motions since the k's have obvious physical significance. This is a big advantage of this calculation, along with the numerical simplicity. We intend to use this method to further investigate the vibrational spectrum of Buckminsterfullerene.

Acknowledgments

We gratefully acknowledge the assistance of Dr. Richard E. Stanton for sharing with us some unpublished research results. We also thank him for useful discussions. This research was supported by the Office of Naval Research and the Air Force Office of Scientific Research (AFSC), United States Air Force, under Contract F49620-86-C-0009. The United States Government is authorized to reproduce and distribute reprints notwithstanding any copyright notation hereon.

References

- 1. H.W. Kroto, J.R. Heath, S.C. O'Brien, R.F. Curl and R.E. Smalley, Nature 318, 162 (1985). For a review of the experimental techniques involved, see A. Kaldor, D.M. Cox, D.J. Trevor and M.R. Zakin, Z. Phys. D 3, 195 (1986).
- 2. Q.L. Zhang, S.C. O'Brien, J.R. Heath, Y. Liu, H.W. Kroto and R.E. Smalley, J. Phys. Chem. 90, 525 (1986).
- 3. A. O'Keefe, M.M. Ross and A.P. Baronavski, Chem. Phys. Lett. <u>130</u>, 17 (1986).
- 4. M.Y. Hahn, E.C. Honea, A. J. Paguia, K.E. Schriver, A.M. Camarena and R.L. Whetten, Chem. Phys. Lett. <u>130</u>, 12 (1986).
- 5. S.C. O'Brien, J.R. Heath, H.W. Kroto, R.F. Curl and R.E. Smalley, Chem. Phys. Lett. 132, 99 (1986).
- 6. M.D. Newton and R.E. Stanton, J. Am. Chem. Soc. 108, 2469 (1986).
- 7. T.G. Schmalz, W.A. Seitz, D.J. Klein and G.E. Hite, Chem. Phys. Lett. 130, 203 (1986).
- 8. S. Satpathy, Chem. Phys. Lett. <u>130</u>, 545 (1986).
- 9. R.C. Haddon, L.E. Brus and K. Raghavachari, Chem. Phys. Lett. 131, 165 (1986).
- 10. P.W. Fowler, Chem. Phys. Lett. <u>131</u>, 444 (1986).
- 11. D.A. Jelski and T.F. George, J. Phys. Chem., submitted.
- 12. E.B. Wilson, J.C. Decius and P.C. Cross, Molecular Vibrations (McGraw-Hill, New York, 1955).
- 13. R.E. Stanton, private communication.
- 14. B.L. Crawford and F.A. Miller, J. Chem. Phys. 17, 249 (1949).

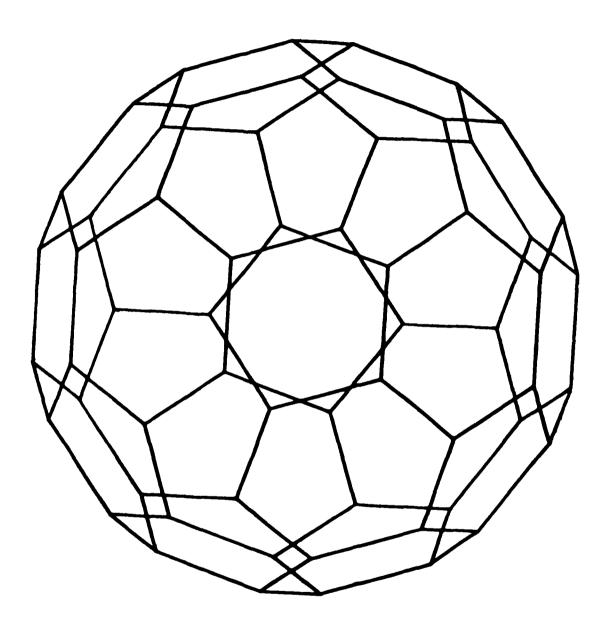
Frequency	Symmetry	Frequency	Symmetry	
0.0	-	272.0	Hg	
354.7	Hu	361.8	F2u	
373.7	Gu	427.9	Hg	
455.8	Gg	491.3	Flu	
491.6	Hu	525.2	F2g	
530.2	Gg	547.6	Ag	
550.9	Flu	552.4	Hg	
566.8	Flg	577.7	Hu	
626.6	F2u	672.7	Gg	
701.9	Gu	726.4	F2g	
755.8	Gu	770.3	Hu	
779.8	Hg	810.9	Flg	
926.6	F2g	958.3	Gu	
1019.0	F2u	1084.3	Au	
1160.0	Hg	1173.5	Gg	
1289.7	Hu	1309.4	F2u	
1374.1	Flu	1398.5	Hg	
1463.8	F2g	1578.5	Hu	
1590.0	Gg	1620.4	Gu	
1627.4	Ag	1655.3	Flu	
1665.1	F2g	1688.2	Hg	
1720.1	F2u	1764.4	Gu	
1765.4	Gg	1830.0	Hu	
1830.7	Hg			

Table 1. Frequencies in cm⁻¹ and their degeneracies for the vibrational modes of Buckminsterfullerene. $k_1 = 1.1 \times 10^6$, $k_2 = 1.0 \times 10^6$, $k_3 = 1.0 \times 10^5$ and $k_4 = 1.0 \times 10^5$ dynes/cm.

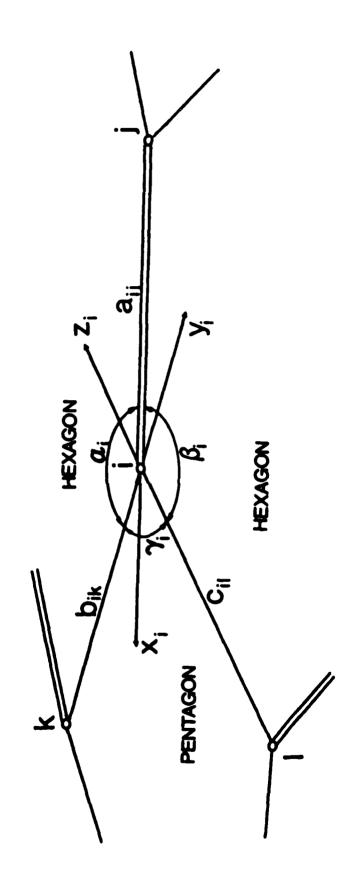
FIGURE CAPTIONS

- Figure 1. The C_{60} cluster, better known as Buckminsterfullerene, is a truncated icosahedron with 20 hexagonal faces and 12 pentagonal faces.
- Figure 2. A diagram of the local coordinate system used in the calculation.

SOUTH SECURITY TOTAL SOUTH SOUTH SECURITY SECURITY SECURITY



EGER MEDICON BEDICON POPOSON OSCOCO BASSASA FACCOCO TASSASAS TASSASAS PASSAS DE PASSAS DE PASSAS DE PASSAS



PASSAMON PARASKY DOMENDON PROCESSES PARA

DOZO BESSESS ZOZOZO BOZOZO BOZOZO KOLOGO BOZOZO

01/1113/86/2

TECHNICAL REPORT DISTRIBUTION LIST, GEN

	No. Copies	•	No. Copies
Office of Naval Research Attn: Code 1113 800 N. Quincy Street Arlington, Virginia 22217-5000	2	Dr. David Young Code 334 NORDA NSTL, Mississippi 39529	1
Dr. Bernard Douda Naval Weapons Support Center Code 50C Crane, Indiana 47522-5050	1	Naval Weapons Center Attn: Dr. Ron Atkins Chemistry Division China Lake, California 93555	1
Naval Civil Engineering Laboratory Attn: Dr. R. W. Drisko, Code L52 Port Hueneme, California 93401	1	Scientific Advisor Commandant of the Marine Corps Code RD-1 Washington, D.C. 20380	1
Defense Technical Information Center Building 5, Cameron Station Alexandria, Virginia 22314	12 high quality	U.S. Army Research Office Attn: CRD-AA-IP P.O. Box 12211 Research Triangle Park, NC 2770	1
DTNSRDC Attn: Dr. H. Singerman Applied Chemistry Division Annapolis, Maryland 21401	1	Mr. John Boyle Materials Branch Naval Ship Engineering Center Philadelphia, Pennsylvania 1911	
Dr. William Tolles Superintendent Chemistry Division, Code 6100 Naval Research Laboratory	1	Naval Ocean Systems Center Attn: Dr. S. Yamamoto Marine Sciences Division San Diego, California 91232	1
Washington, D.C. 20375-5000	. •	Dr. David L. Nelson Chemistry Division Office of Naval Research 800 North Quincy Street Arlington, Virginia 22217	1

Dr. J. E. Jensen Hughes Research Laboratory 3011 Malibu Canyon Road Malibu. California 90265

Dr. J. H. Weaver
Department of Chemical Engineering
and Materials Science
University of Minnesota
Minneapolis, Minnesota 55455

Dr. A. Reisman

Microelectronics Center of North Carolina Dr. D. Dilella

Research Triangle Park, North Carolina Chemistry Depa

27709 George Washing

Dr. M. Grunze
Laboratory for Surface Science and
Technology
University of Maine
Orono, Maine 04469

Dr. J. Butler Naval Research Laboratory Code 6115 Washington D.C. 20375-5000

Dr. L. Interante Chemistry Department Rensselaer Polytechnic Institute Troy, New York 12181

Or. Irvin Heard Chemistry and Physics Department Lincoln University Lincoln University, Pennsylvania 19352

Dr. K.J. Klaubunde Department of Chemistry Kansas State University Manhattan, Kansas 66506 Dr. C. B. Harris Department of Chemistry University of California Berkeley, California 94720

Dr. F. Kutzler
Department of Chemistry
Box 5055
Tennessee Technological University
Cookesville, Tennessee 38501

Dr. D. Dilella Chemistry Department George Washington University Washington D.C. 20052

Dr. R. Reeves Chemistry Department Renssaeler Polytechnic Institute Troy, New York 12181

Dr. Steven M. George Stanford University Department of Chemistry Stanford, CA 94305

Dr. Mark Johnson Yale University Department of Chemistry New Haven, CT 06511-8118

Dr. W. Knauer Hughes Research Laboratory 3011 Malibu Canyon Road Malibu, California 90265

Dr. G. A. Somorjai Department of Chemistry University of California Berkeley, California 94720

Dr. J. Murday
Naval Research Laboratory
Code 6170
Washington, D.C. 20375-5000

Dr. J. B. Hudson Materials Division Rensselaer Polytechnic Institute Troy, New York 12181

Dr. Theodore E. Madey Surface Chemistry Section Department of Commerce National Bureau of Standards Washington, D.C. 20234

Dr. J. E. Demuth
IBM Corporation
Thomas J. Watson Research Center
P.O. Box 218
Yorktown Heights, New York 10598

Dr. M. G. Lagally
Department of Metallurgical
and Mining Engineering
University of Wisconsin
Madison, Wisconsin 53706

Dr. R. P. Van Duyne Chemistry Department Northwestern University Evanston, Illinois 60637

Dr. J. M. White Department of Chemistry University of Texas Austin, Texas 78712

Dr. D. E. Harrison Department of Physics Naval Postgraduate School Monterey, California 93940 Dr. R. L. Park
Director, Center of Materials
Research
University of Maryland
College Park, Maryland 20742

Dr. W. T. Peria Electrical Engineering Department University of Minnesota Minneapolis, Minnesota 55455

Or. Keith H. Johnson
Department of Metallurgy and
Materials Science
Massachusetts Institute of Technology
Cambridge, Massachusetts 02139

Dr. S. Sibener
Department of Chemistry
James Franck Institute
5640 Ellis Avenue
Chicago, Illinois 60637

Dr. Arnold Green Quantum Surface Dynamics Branch Code 3817 Naval Weapons Center China Lake, California 93555

Dr. A. Wold Department of Chemistry Brown University Providence, Rhode Island 02912

Dr. S. L. Bernasek Department of Chemistry Princeton University Princeton, New Jersey 08544

Dr. W. Kohn
Department of Physics
University of California, San Diego
La Jolla, California 92037

Dr. F. Carter Code 6170 Naval Research Laboratory Washington, D.C. 20375-5000

Or. Richard Colton Code 6170 Naval Research Laboratory Washington, D.C. 20375-5000

Dr. Dan Pierce National Bureau of Standards Optical Physics Division Washington, D.C. 20234

Dr. R. Stanley Williams
Department of Chemistry
University of California
Los Angeles, California 90024

Dr. R. P. Messmer Materials Characterization Lab. General Electric Company Schenectady, New York 22217

Or. Robert Gomer Department of Chemistry James Franck Institute 5640 Ellis Avenue Chicago, Illinois 60637

Dr. Ronald Lee R301 Naval Surface Weapons Center White Oak Silver Spring, Maryland 20910

Dr. Paul Schoen Code 6190 Naval Research Laboratory Washington, D.C. 20375-5000 Dr. John T. Yates Department of Chemistry University of Pittsburgh Pittsburgh, Pennsylvania 15260

Dr. Richard Greene Code 5230 Naval Research Laboratory Washington, D.C. 20375-5000

Dr. L. Kesmodel
Department of Physics
Indiana University
Bloomington, Indiana 47403

Dr. K. C. Janda University of Pittsburg Chemistry Building Pittsburg, PA 15260

Dr. E. A. Irene
Department of Chemistry
University of North Carolina
Chapel Hill, North Carolina 27514

Dr. Adam Heller Bell Laboratories Murray Hill, New Jersey 07974

Dr. Martin Fleischmann Department of Chemistry University of Southampton Southampton 509 5MH UNITED KINGDOM

Dr. H. Tachikawa Chemistry Department Jackson State University Jackson, Mississippi 39217 SPERT TERSOSSIES ("TERSOSSIES PRODUCES PROCESSIES PROCESSIES PROCESSIES

Dr. John W. Wilkins Cornell University Laboratory of Atomic and Solid State Physics Ithaca, New York 14853

Dr. R. G. Wallis Department of Physics University of California Irvine, California 92664

Dr. D. Ramaker Chemistry Department George Washington University Washington, D.C. 20052

Dr. J. C. Hemminger . Chemistry Department University of California Irvine, California 92717

Dr. T. F. George Chemistry Department University of Rochester Rochester, New York 14627

Or. G. Rubloff
IBM
Thomas J. Watson Research Center
P.O. Box 218
Yorktown Heights, New York 10598

Dr. Horia Metiu Chemistry Department University of California Santa Barbara, California 93106

Dr. W. Goddard
Department of Chemistry and Chemical
Engineering
California Institute of Technology
Pasadena, California 91125

Dr. P. Hansma Department of Physics University of California Santa Barbara, California 93106

Dr. J. Baldeschwieler
Department of Chemistry and
Chemical Engineering
California Institute of Technology
Pasadena, California 91125

Dr. J. T. Keiser Department of Chemistry University of Richmond Richmond, Virginia 23173

Dr. R. W. Plummer
Department of Physics
University of Pennsylvania
Philadelphia, Pennsylvania 19104

Dr. E. Yeager Department of Chemistry Case Western Reserve University Cleveland, Ohio 41106

Dr. N. Winograd
Department of Chemistry
Pennsylvania State University
University Park, Pennsylvania 16802

Dr. Roald Hoffmann Department of Chemistry Cornell University Ithaca, New York 14853

Dr. A. Steck1
Department of Electrical and
Systems Engineering
Rensselaer Polytechnic Institute
Troy, NewYork 12181

Dr. G.H. Morrison Department of Chemistry Cornell University Ithaca, New York 14853

4-87

1010